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Department:
Basic Education
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NATIONAL SENIOR CERTIFICATE

GRADE 12

PHYSICAL SCIENCES: CHEMISTRY (P2)

NOVEMBER 2019

MARKS: 150

TIME: 3 hours

This question paper consists of 14 pages and 4 data sheets.

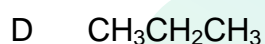
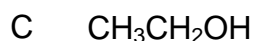
INSTRUCTIONS AND INFORMATION

1. Write your centre number and examination number in the appropriate spaces on the ANSWER BOOK.
2. This question paper consists of TEN questions. Answer ALL the questions in the ANSWER BOOK.
3. Start EACH question on a NEW page in the ANSWER BOOK.
4. Number the answers correctly according to the numbering system used in this question paper.
5. Leave ONE line between two subquestions, e.g. between QUESTION 2.1 and QUESTION 2.2.
6. You may use a non-programmable calculator.
7. You may use appropriate mathematical instruments.
8. You are advised to use the attached DATA SHEETS.
9. Show ALL formulae and substitutions in ALL calculations.
10. Round off your FINAL numerical answers to a minimum of TWO decimal places.
11. Give brief motivations, discussions, etc. where required.
12. Write neatly and legibly.

QUESTION 1: MULTIPLE-CHOICE QUESTIONS

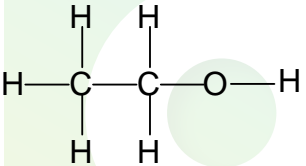
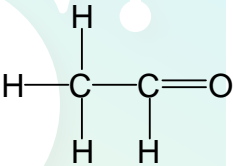
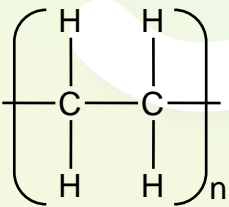
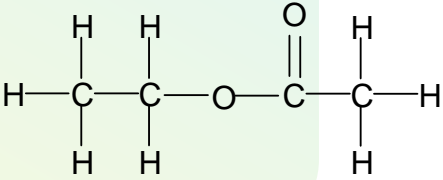
Four options are provided as possible answers to the following questions. Choose the answer and write only the letter (A–D) next to the question numbers (1.1 to 1.10) in the ANSWER BOOK, e.g. 1.11 E.

1.1 Which ONE of the following compounds has the HIGHEST vapour pressure?



(2)

1.2 Which ONE of the formulae below represents the product of a POLYMERISATION reaction?

A		B	
C		D	

(2)

1.3 Which ONE of the following combinations are BOTH UNSATURATED HYDROCARBONS?

A Ethane and ethene

B Ethene and ethyne

C Ethane and ethanol

D Ethanoic acid and ethene

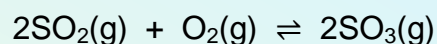
(2)

- 1.4 Which ONE of the following sets of values for activation energy (E_a) and heat of reaction (ΔH) is possible for a reaction?

	ACTIVATION ENERGY (E_a) ($\text{kJ}\cdot\text{mol}^{-1}$)	HEAT OF REACTION (ΔH) ($\text{kJ}\cdot\text{mol}^{-1}$)
A	100	+100
B	50	+100
C	50	+50
D	100	-50

(2)

- 1.5 Consider the following balanced equation for a system at equilibrium:

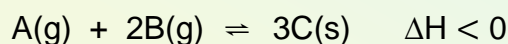


How will the addition of a catalyst to the equilibrium mixture affect the YIELD and REACTION RATE?

	YIELD	REACTION RATE
A	Increases	Increases
B	Remains the same	Remains the same
C	Remains the same	Increases
D	Decreases	Increases

(2)

- 1.6 A hypothetical reaction reaches equilibrium at a certain temperature in a closed container according to the following balanced equation:



Which ONE of the following changes to the equilibrium conditions will result in an INCREASE in the equilibrium constant, K_c ?

- A Increase in temperature
- B Decrease in temperature
- C Increase in pressure at constant temperature
- D Decrease in pressure at constant temperature

(2)

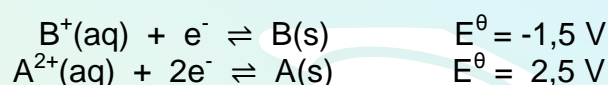
- 1.7 A hydrochloric acid solution, $\text{HCl}(\text{aq})$, and an acetic acid solution, $\text{CH}_3\text{COOH}(\text{aq})$, of EQUAL CONCENTRATIONS are compared.

How do the $\text{H}_3\text{O}^+(\text{aq})$ concentration of $\text{HCl}(\text{aq})$ and the pH of $\text{HCl}(\text{aq})$ compare to that of $\text{CH}_3\text{COOH}(\text{aq})$?

	$[\text{H}_3\text{O}^+]$ of $\text{HCl}(\text{aq})$	pH of $\text{HCl}(\text{aq})$
A	Higher than	Higher than
B	Higher than	Lower than
C	Equal to	Equal to
D	Higher than	Equal to

(2)

- 1.8 Two hypothetical half-reactions and their respective reduction potentials are shown below:



A galvanic cell is set up using the above substances.

Which ONE of the following statements is CORRECT for this galvanic cell?

- A $\text{B}(\text{s})$ is the reducing agent.
- B $\text{A}(\text{s})$ is the oxidising agent.
- C The mass of $\text{B}(\text{s})$ will increase.
- D The mass of $\text{A}(\text{s})$ will decrease.

(2)

- 1.9 In an electrolytic cell ...

- A the anode is the positive electrode.
- B oxidation takes place at the cathode.
- C electrons flow from the cathode to the anode.
- D the mass of the anode increases.

(2)

- 1.10 Which ONE of the following is used as a catalyst in the Ostwald process?

- A Iron
- B Copper
- C Platinum
- D Vanadium (V) oxide

(2)
[20]

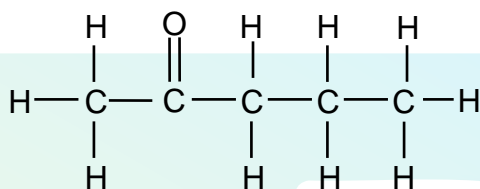
QUESTION 2 (Start on a new page.)

2.1 The IUPAC name of an organic compound is 4,4-dimethylpent-2-yne.

2.1.1 Write down the GENERAL FORMULA of the homologous series to which this compound belongs. (1)

2.1.2 Write down the STRUCTURAL formula of this compound. (3)

2.2 The organic compound below has one positional isomer and one functional isomer.



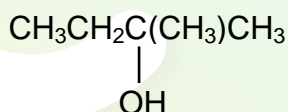
2.2.1 Define the term *positional isomer*. (2)

For this compound, write down the:

2.2.2 IUPAC name of its POSITIONAL isomer (2)

2.2.3 Structural formula of its FUNCTIONAL isomer (2)

2.3 Consider the condensed structural formula of an organic compound below.



2.3.1 Is this a PRIMARY, SECONDARY or TERTIARY alcohol? Give a reason for the answer. (2)

2.3.2 Write down the IUPAC name of the above compound. (2)

2.3.3 Write down the IUPAC name of the MAJOR ORGANIC PRODUCT formed when this compound undergoes an elimination reaction. (2)

[16]

QUESTION 3 (Start on a new page.)

The boiling points of five organic compounds (**P**, **Q**, **R**, **S** and **T**) are studied.

COMPOUND	IUPAC NAME
P	Pentanal
Q	2,2-dimethylbutane
R	3-methylpentane
S	Hexane
T	Pentan-1-ol

3.1 Define the term *boiling point*. (2)

The boiling points of compounds **Q**, **R** and **S** are compared.

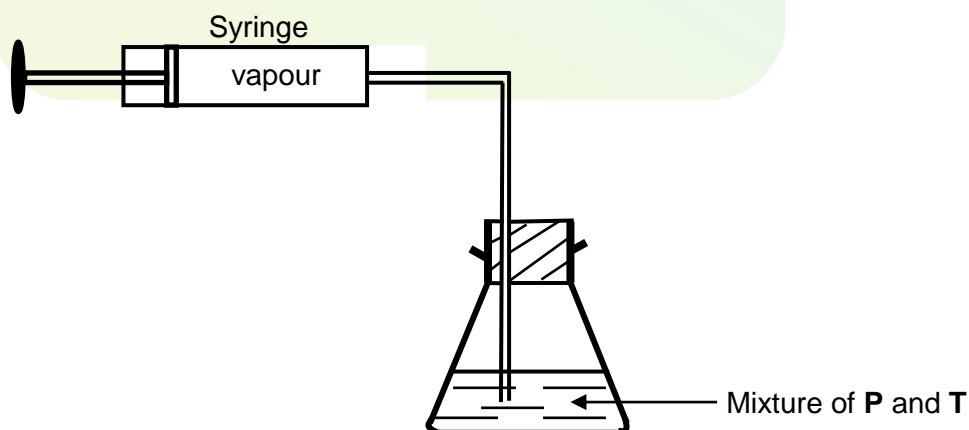
3.2 Give a reason why this is a fair comparison. (1)

The boiling points of **Q**, **R** and **S** are given below (NOT necessarily in the correct order).

55 °C	49,7 °C	68 °C
-------	---------	-------

3.3 Which ONE of the three boiling points is most likely the boiling point of compound **R**? Explain the answer. (4)

3.4 A mixture of equal amounts of **P** and **T** is placed in a flask and heated to a temperature below their boiling points. Assume that no reaction or condensation takes place. The vapour produced is collected in a syringe.

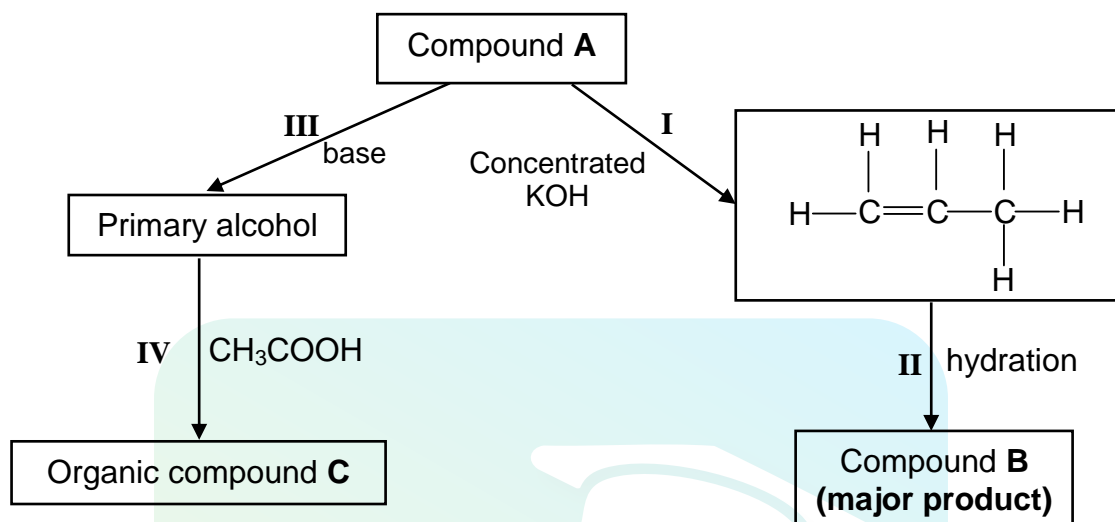


3.4.1 Which compound (**P** or **T**) will be present in a greater amount in the SYRINGE? (2)

3.4.2 Explain the answer to QUESTION 3.4.1 by referring to the TYPES and STRENGTHS of intermolecular forces. (3)
[12]

QUESTION 4 (Start on a new page.)

The flow diagram below shows how compound **A** can be used to prepare other organic compounds. The numbers **I**, **II**, **III** and **IV** represent different organic reactions.



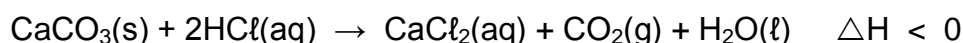
Use the information in the flow diagram to answer the following questions.

- 4.1 Name the homologous series to which compound **A** belongs. (1)
- 4.2 Write down the TYPE of reaction represented by:
 - 4.2.1 **I** (1)
 - 4.2.2 **III** (1)
 - 4.2.3 **IV** (1)
- 4.3 Consider **reaction III**. Write down the:
 - 4.3.1 TWO reaction conditions for this reaction (2)
 - 4.3.2 IUPAC name of the primary alcohol that is formed (2)
- 4.4 Draw the STRUCTURAL FORMULA for compound **B**. (2)
- 4.5 Consider **reaction IV**. Write down the:
 - 4.5.1 Structural formula of organic compound **C** (2)
 - 4.5.2 NAME or FORMULA of the catalyst that is used (1)

[13]

QUESTION 5 (Start on a new page.)

The calcium carbonate (CaCO_3) in antacid tablets reacts with dilute hydrochloric acid (HCl) according to the following balanced equation:



- 5.1 Is the above reaction EXOTHERMIC or ENDOTHERMIC? Give a reason for the answer. (2)

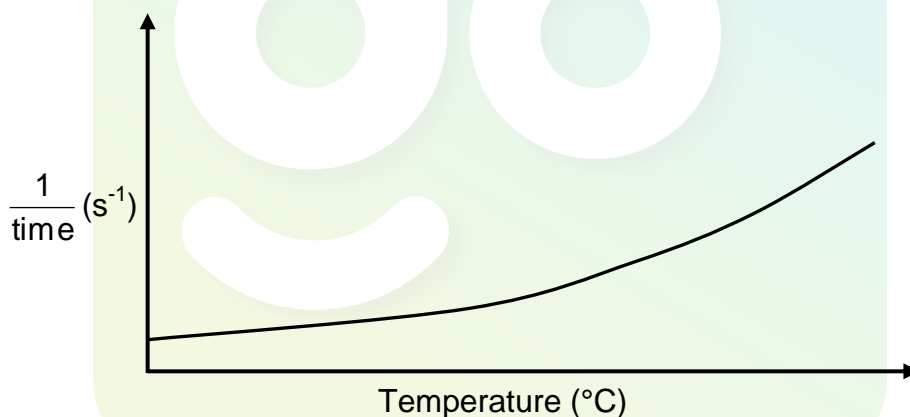
An antacid tablet of mass 2 g is placed in $\text{HCl}(\text{aq})$. After 30 s the mass of the tablet was found to be 0,25 g.

- 5.2 Calculate the average rate (in $\text{g}\cdot\text{s}^{-1}$) of the above reaction. (3)

The antacid tablet contains 40% calcium carbonate. Another antacid tablet of mass 2 g is allowed to react completely with $\text{HCl}(\text{aq})$.

- 5.3 Calculate the volume of carbon dioxide, $\text{CO}_2(\text{g})$ that will be collected at STP. Assume that all the $\text{CO}_2(\text{g})$ produced is from the calcium carbonate. (5)

The reaction rate of similar antacid tablets with excess $\text{HCl}(\text{aq})$ of concentration $0,1 \text{ mol}\cdot\text{dm}^{-3}$ at DIFFERENT TEMPERATURES is measured. The graph below was obtained.



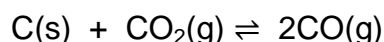
Use the information in the graph to answer the following questions.

- 5.4 Write down ONE controlled variable for this investigation. (1)
- 5.5 Write down a conclusion that can be made from the graph. (2)
- 5.6 Use the collision theory to fully explain the answer to QUESTION 5.5 (3)
- 5.7 Redraw the graph above in the ANSWER BOOK. On the same set of axes, sketch the curve that will be obtained if $\text{HCl}(\text{aq})$ of concentration $0,2 \text{ mol}\cdot\text{dm}^{-3}$ is now used. Label this curve Y. (2)

[18]

QUESTION 6 (Start on a new page.)

Initially 60,8 g pure carbon dioxide, $\text{CO}_2(\text{g})$, is reacted with carbon, $\text{C}(\text{s})$, in a sealed container of volume 3 dm^3 . The reaction reaches equilibrium at temperature **T** according to the following balanced equation:



6.1 Define the term *chemical equilibrium*. (2)

6.2 At equilibrium it is found that the concentration of the carbon dioxide is $0,054 \text{ mol} \cdot \text{dm}^{-3}$.

Calculate the:

6.2.1 Equilibrium constant, K_c , for this reaction at temperature **T** (7)

6.2.2 Minimum mass of $\text{C}(\text{s})$ that must be present in the container to obtain this equilibrium (3)

6.3 How will EACH of the following changes affect the AMOUNT of $\text{CO}(\text{g})$ at equilibrium?

Choose from INCREASES, DECREASES or REMAINS THE SAME.

6.3.1 More carbon is added to the container (1)

6.3.2 The pressure is increased by reducing the volume of the container at constant temperature.
Use Le Chatelier's principle to explain the answer. (3)

6.4 The table below shows the percentages of $\text{CO}_2(\text{g})$ and $\text{CO}(\text{g})$ in the container at different temperatures.

TEMPERATURE ($^{\circ}\text{C}$)	% $\text{CO}_2(\text{g})$	% $\text{CO}(\text{g})$
827	6,23	93,77
950	1,32	98,68
1 050	0,37	99,63
1 200	0,06	99,94

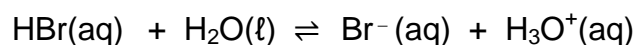
6.4.1 Is the reaction EXOTHERMIC or ENDOTHERMIC?
Refer to the data in the table and explain the answer. (3)

6.4.2 Use the information in the table to determine temperature **T**.
Show clearly how you arrived at the answer. (3)

[22]

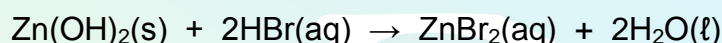
QUESTION 7 (Start on a new page.)

A hydrogen bromide solution, HBr(aq) , reacts with water according to the following balanced chemical equation:



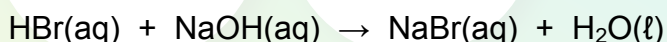
The K_a value of HBr(aq) at 25°C is 1×10^9 .

- 7.1 Is hydrogen bromide a STRONG ACID or a WEAK ACID? Give a reason for the answer. (2)
- 7.2 Write down the FORMULAE of the TWO bases in the above reaction. (2)
- 7.3 HBr(aq) reacts with $\text{Zn(OH)}_2(\text{s})$ according to the following balanced equation:



An unknown quantity of $\text{Zn(OH)}_2(\text{s})$ is reacted with 90 cm^3 of HBr(aq) in a flask. (Assume that the volume of the solution does not change during the reaction.)

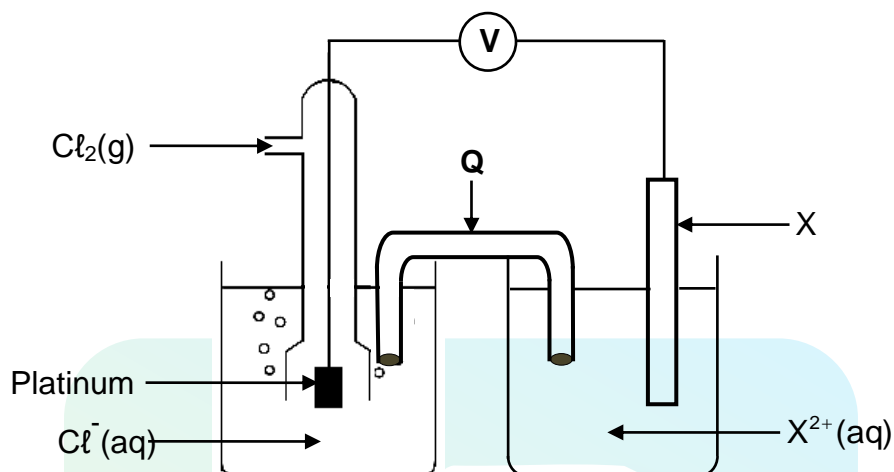
The EXCESS HBr(aq) is then neutralised by $16,5 \text{ cm}^3$ of NaOH(aq) of concentration $0,5 \text{ mol}\cdot\text{dm}^{-3}$. The balanced equation for the reaction is:



- 7.3.1 Calculate the pH of the HBr solution remaining in the flask AFTER the reaction with $\text{Zn(OH)}_2(\text{s})$. (7)
- 7.3.2 Calculate the mass of $\text{Zn(OH)}_2(\text{s})$ INITIALLY present in the flask if the initial concentration of HBr(aq) was $0,45 \text{ mol}\cdot\text{dm}^{-3}$. (6)
- [17]**

QUESTION 8 (Start on a new page.)

A standard electrochemical cell is set up using two standard half-cells, as shown in the diagram below.



8.1 State the energy conversion that takes place in this cell. (1)

8.2 What is the function of component **Q**? (1)

X is a metal. A voltmeter connected across the cell initially registers 1,49 V.

8.3 Use a calculation to identify metal **X**. (5)

8.4 Write down the NAME or FORMULA of the reducing agent. (1)

8.5 The reading on the voltmeter becomes ZERO after this cell operates for several hours.

8.5.1 Give a reason for this reading by referring to the rates of oxidation and reduction half-reactions taking place in the cell. (1)

A silver nitrate solution, $\text{AgNO}_3(\text{aq})$, is NOW added to the chlorine half-cell and a precipitate forms.

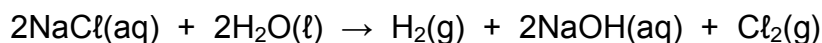
8.5.2 How will the reading on the voltmeter be affected?
(Choose from INCREASES, DECREASES or REMAINS the same) (1)

8.5.3 Use Le Chatelier's principle to explain the answer to QUESTION 8.5.2. (2)
[12]

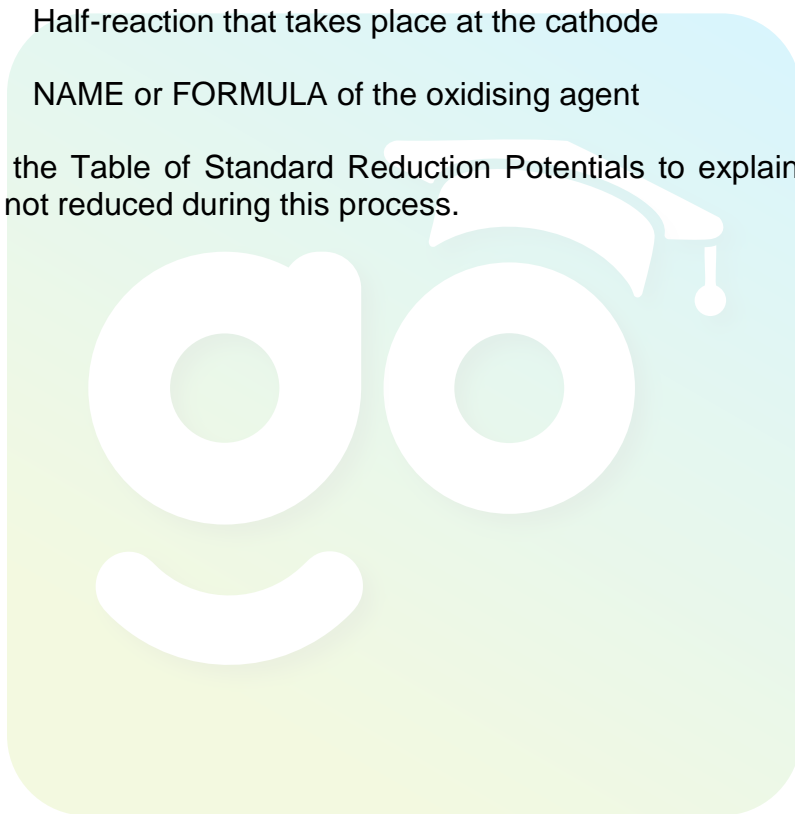
QUESTION 9 (Start on a new page.)

Chlorine is produced industrially by the electrolysis of a concentrated sodium chloride solution, $\text{NaCl}(\text{aq})$.

The balanced equation for the net (overall) cell reaction is as follows:

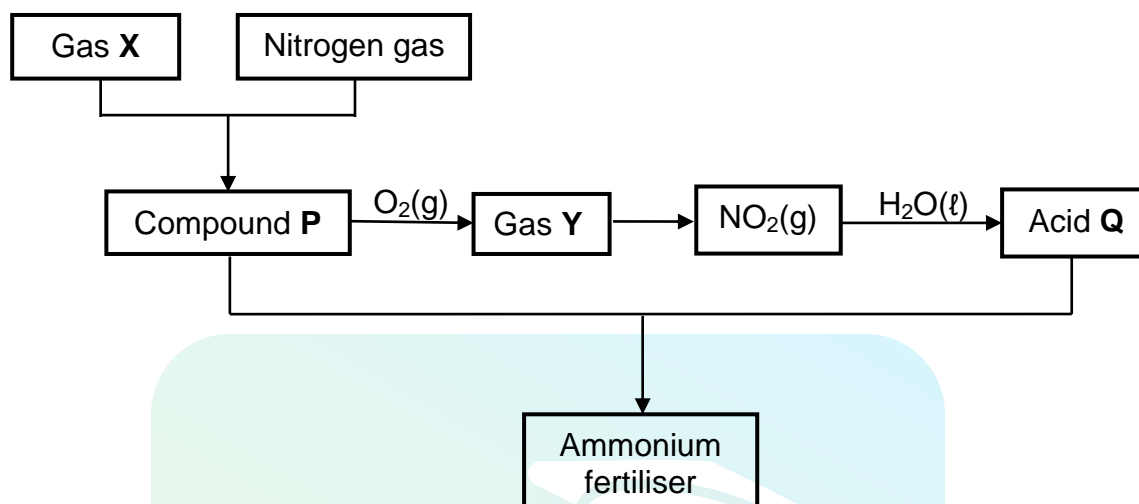


- 9.1 Define the term *electrolysis*. (2)
- 9.2 For the above reaction, write down the:
- 9.2.1 Half-reaction that takes place at the cathode (2)
- 9.2.2 NAME or FORMULA of the oxidising agent (1)
- 9.3 Refer to the Table of Standard Reduction Potentials to explain why sodium ions are not reduced during this process. (3)
- [8]**



QUESTION 10 (Start on a new page.)

The flow diagram below shows the processes involved in the industrial preparation of an ammonium fertiliser.



10.1 Write down the NAME or FORMULA of:

10.1.1 Gas **X** (1)

10.1.2 Gas **Y** (1)

10.1.3 Acid **Q** (1)

10.2 Write down the:

10.2.1 TYPE of chemical reaction that converts compound **P** into gas **Y** (1)

10.2.2 Balanced equation for the reaction between compound **P** and acid **Q** (3)

10.3 Two separate bags of fertilisers are labelled as follows:

BAG A
1 : 3 : 4 (21)
50 kg

BAG B
1 : 3 : 4 (27)
40 kg

10.3.1 What do the numbers (21) and (27) on the labels represent? (1)

10.3.2 Determine, by means of calculations, which bag (**A** or **B**) contains a greater mass of phosphorous. (4)
[12]

TOTAL: 150

**DATA FOR PHYSICAL SCIENCES GRADE 12
PAPER 2 (CHEMISTRY)**

**GEGEWENS VIR FISIIESE WETENSKAPPE GRAAD 12
VRAESTEL 2 (CHEMIE)**

TABLE 1: PHYSICAL CONSTANTS/TABEL 1: FISIIESE KONSTANTES

NAME/NAAM	SYMBOL/SIMBOOL	VALUE/WAARDE
Standard pressure <i>Standaarddruk</i>	p^{θ}	$1,013 \times 10^5 \text{ Pa}$
Molar gas volume at STP <i>Molêre gasvolume by STD</i>	V_m	$22,4 \text{ dm}^3 \cdot \text{mol}^{-1}$
Standard temperature <i>Standaardtemperatuur</i>	T^{θ}	273 K
Charge on electron <i>Lading op elektron</i>	e	$-1,6 \times 10^{-19} \text{ C}$
Avogadro's constant <i>Avogadro-konstante</i>	N_A	$6,02 \times 10^{23} \text{ mol}^{-1}$

TABLE 2: FORMULAE/TABEL 2: FORMULES

$n = \frac{m}{M}$	$n = \frac{N}{N_A}$
$c = \frac{n}{V}$ or/of $c = \frac{m}{MV}$	$n = \frac{V}{V_m}$
$\frac{c_a v_a}{c_b v_b} = \frac{n_a}{n_b}$	$\text{pH} = -\log[\text{H}_3\text{O}^+]$
$K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1 \times 10^{-14} \text{ at/by } 298 \text{ K}$	
$E_{\text{cell}}^{\theta} = E_{\text{cathode}}^{\theta} - E_{\text{anode}}^{\theta} / E_{\text{sel}}^{\theta} = E_{\text{katode}}^{\theta} - E_{\text{anode}}^{\theta}$ or/of $E_{\text{cell}}^{\theta} = E_{\text{reduction}}^{\theta} - E_{\text{oxidation}}^{\theta} / E_{\text{sel}}^{\theta} = E_{\text{reduksie}}^{\theta} - E_{\text{oksidasie}}^{\theta}$ or/of $E_{\text{cell}}^{\theta} = E_{\text{oxidisingagent}}^{\theta} - E_{\text{reducingagent}}^{\theta} / E_{\text{sel}}^{\theta} = E_{\text{oksideermiddel}}^{\theta} - E_{\text{reduseermiddel}}^{\theta}$	

TABLE 3: THE PERIODIC TABLE OF ELEMENTS
TABEL 3: DIE PERIODIEKE TABEL VAN ELEMENTE

1 (I)	2 (II)	3	4	5	6	7	8	9	10	11	12	13 (III)	14 (IV)	15 (V)	16 (VI)	17 (VII)	18 (VIII)
1 2,1 H 1																	2 He 4
3 1,0 Li 7	4 1,5 Be 9											5 2,0 B 11	6 2,5 C 12	7 3,0 N 14	8 3,5 O 16	9 4,0 F 19	10 Ne 20
11 0,9 Na 23	12 1,2 Mg 24											13 1,5 Al 27	14 1,8 Si 28	15 2,1 P 31	16 2,5 S 32	17 3,0 Cl 35,5	18 Ar 40
19 0,8 K 39	20 1,0 Ca 40	21 1,3 Sc 45	22 1,5 Ti 48	23 1,6 V 51	24 1,6 Cr 52	25 1,5 Mn 55	26 1,8 Fe 56	27 1,8 Co 59	28 1,8 Ni 59	29 1,9 Cu 63,5	30 1,6 Zn 65	31 1,6 Ga 70	32 1,8 Ge 73	33 2,0 As 75	34 2,4 Se 79	35 2,8 Br 80	36 Kr 84
37 0,8 Rb 86	38 1,0 Sr 88	39 1,2 Y 89	40 1,4 Zr 91	41 Nb 92	42 1,8 Mo 96	43 1,9 Tc	44 2,2 Ru 101	45 2,2 Rh 103	46 2,2 Pd 106	47 1,9 Ag 108	48 1,7 Cd 112	49 1,7 In 115	50 1,8 Sn 119	51 1,9 Sb 122	52 2,1 Te 128	53 2,5 I 127	54 Xe 131
55 0,7 Cs 133	56 0,9 Ba 137	57 La 139	72 1,6 Hf 179	73 Ta 181	74 W 184	75 Re 186	76 Os 190	77 Ir 192	78 Pt 195	79 Au 197	80 Hg 201	81 1,8 Tl 204	82 1,8 Pb 207	83 1,9 Bi 209	84 2,0 Po	85 2,5 At	86 Rn
87 0,7 Fr	88 0,9 Ra 226	89 Ac															
			58 Ce 140	59 Pr 141	60 Nd 144	61 Pm	62 Sm 150	63 Eu 152	64 Gd 157	65 Tb 159	66 Dy 163	67 Ho 165	68 Er 167	69 Tm 169	70 Yb 173	71 Lu 175	
			90 Th 232	91 Pa	92 U 238	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr	

KEY/SLEUTEL

Atomic number
Atoomgetal

Electronegativity
Elektronegatiwiteit

Symbol
Simbool

Approximate relative atomic mass
Benaderde relatiewe atoommassa

29
Cu
63,5

KEY/SLEUTEL

Atomic number
Atoomgetal

Electronegativity
Elektronegatiwiteit

Symbol
Simbool

Approximate relative atomic mass
Benaderde relatiewe atoommassa

TABLE 4A: STANDARD REDUCTION POTENTIALS
TABEL 4A: STANDAARD-REDUKSIEPOTENSIALE

Half-reactions/ <i>Halfreaksies</i>	E^{θ} (V)
$F_2(g) + 2e^- \rightleftharpoons 2F^-$	+ 2,87
$Co^{3+} + e^- \rightleftharpoons Co^{2+}$	+ 1,81
$H_2O_2 + 2H^+ + 2e^- \rightleftharpoons 2H_2O$	+1,77
$MnO_4^- + 8H^+ + 5e^- \rightleftharpoons Mn^{2+} + 4H_2O$	+ 1,51
$Cl_2(g) + 2e^- \rightleftharpoons 2Cl^-$	+ 1,36
$Cr_2O_7^{2-} + 14H^+ + 6e^- \rightleftharpoons 2Cr^{3+} + 7H_2O$	+ 1,33
$O_2(g) + 4H^+ + 4e^- \rightleftharpoons 2H_2O$	+ 1,23
$MnO_2 + 4H^+ + 2e^- \rightleftharpoons Mn^{2+} + 2H_2O$	+ 1,23
$Pt^{2+} + 2e^- \rightleftharpoons Pt$	+ 1,20
$Br_2(l) + 2e^- \rightleftharpoons 2Br^-$	+ 1,07
$NO_3^- + 4H^+ + 3e^- \rightleftharpoons NO(g) + 2H_2O$	+ 0,96
$Hg^{2+} + 2e^- \rightleftharpoons Hg(l)$	+ 0,85
$Ag^+ + e^- \rightleftharpoons Ag$	+ 0,80
$NO_3^- + 2H^+ + e^- \rightleftharpoons NO_2(g) + H_2O$	+ 0,80
$Fe^{3+} + e^- \rightleftharpoons Fe^{2+}$	+ 0,77
$O_2(g) + 2H^+ + 2e^- \rightleftharpoons H_2O_2$	+ 0,68
$I_2 + 2e^- \rightleftharpoons 2I^-$	+ 0,54
$Cu^+ + e^- \rightleftharpoons Cu$	+ 0,52
$SO_2 + 4H^+ + 4e^- \rightleftharpoons S + 2H_2O$	+ 0,45
$2H_2O + O_2 + 4e^- \rightleftharpoons 4OH^-$	+ 0,40
$Cu^{2+} + 2e^- \rightleftharpoons Cu$	+ 0,34
$SO_4^{2-} + 4H^+ + 2e^- \rightleftharpoons SO_2(g) + 2H_2O$	+ 0,17
$Cu^{2+} + e^- \rightleftharpoons Cu^+$	+ 0,16
$Sn^{4+} + 2e^- \rightleftharpoons Sn^{2+}$	+ 0,15
$S + 2H^+ + 2e^- \rightleftharpoons H_2S(g)$	+ 0,14
$2H^+ + 2e^- \rightleftharpoons H_2(g)$	0,00
$Fe^{3+} + 3e^- \rightleftharpoons Fe$	- 0,06
$Pb^{2+} + 2e^- \rightleftharpoons Pb$	- 0,13
$Sn^{2+} + 2e^- \rightleftharpoons Sn$	- 0,14
$Ni^{2+} + 2e^- \rightleftharpoons Ni$	- 0,27
$Co^{2+} + 2e^- \rightleftharpoons Co$	- 0,28
$Cd^{2+} + 2e^- \rightleftharpoons Cd$	- 0,40
$Cr^{3+} + e^- \rightleftharpoons Cr^{2+}$	- 0,41
$Fe^{2+} + 2e^- \rightleftharpoons Fe$	- 0,44
$Cr^{3+} + 3e^- \rightleftharpoons Cr$	- 0,74
$Zn^{2+} + 2e^- \rightleftharpoons Zn$	- 0,76
$2H_2O + 2e^- \rightleftharpoons H_2(g) + 2OH^-$	- 0,83
$Cr^{2+} + 2e^- \rightleftharpoons Cr$	- 0,91
$Mn^{2+} + 2e^- \rightleftharpoons Mn$	- 1,18
$Al^{3+} + 3e^- \rightleftharpoons Al$	- 1,66
$Mg^{2+} + 2e^- \rightleftharpoons Mg$	- 2,36
$Na^+ + e^- \rightleftharpoons Na$	- 2,71
$Ca^{2+} + 2e^- \rightleftharpoons Ca$	- 2,87
$Sr^{2+} + 2e^- \rightleftharpoons Sr$	- 2,89
$Ba^{2+} + 2e^- \rightleftharpoons Ba$	- 2,90
$Cs^+ + e^- \rightleftharpoons Cs$	- 2,92
$K^+ + e^- \rightleftharpoons K$	- 2,93
$Li^+ + e^- \rightleftharpoons Li$	- 3,05

Increasing oxidising ability/*Toenemende oksiderende vermoë*Increasing reducing ability/*Toenemende reducerende vermoë*

TABLE 4B: STANDARD REDUCTION POTENTIALS
TABEL 4B: STANDAARD-REDUKSIEPOTENSIALE

Half-reactions/ <i>Halfreaksies</i>	E^{θ} (V)
$\text{Li}^{+} + \text{e}^{-} \rightleftharpoons \text{Li}$	-3,05
$\text{K}^{+} + \text{e}^{-} \rightleftharpoons \text{K}$	-2,93
$\text{Cs}^{+} + \text{e}^{-} \rightleftharpoons \text{Cs}$	-2,92
$\text{Ba}^{2+} + 2\text{e}^{-} \rightleftharpoons \text{Ba}$	-2,90
$\text{Sr}^{2+} + 2\text{e}^{-} \rightleftharpoons \text{Sr}$	-2,89
$\text{Ca}^{2+} + 2\text{e}^{-} \rightleftharpoons \text{Ca}$	-2,87
$\text{Na}^{+} + \text{e}^{-} \rightleftharpoons \text{Na}$	-2,71
$\text{Mg}^{2+} + 2\text{e}^{-} \rightleftharpoons \text{Mg}$	-2,36
$\text{Al}^{3+} + 3\text{e}^{-} \rightleftharpoons \text{Al}$	-1,66
$\text{Mn}^{2+} + 2\text{e}^{-} \rightleftharpoons \text{Mn}$	-1,18
$\text{Cr}^{2+} + 2\text{e}^{-} \rightleftharpoons \text{Cr}$	-0,91
$2\text{H}_2\text{O} + 2\text{e}^{-} \rightleftharpoons \text{H}_2(\text{g}) + 2\text{OH}^{-}$	-0,83
$\text{Zn}^{2+} + 2\text{e}^{-} \rightleftharpoons \text{Zn}$	-0,76
$\text{Cr}^{3+} + 3\text{e}^{-} \rightleftharpoons \text{Cr}$	-0,74
$\text{Fe}^{2+} + 2\text{e}^{-} \rightleftharpoons \text{Fe}$	-0,44
$\text{Cr}^{3+} + \text{e}^{-} \rightleftharpoons \text{Cr}^{2+}$	-0,41
$\text{Cd}^{2+} + 2\text{e}^{-} \rightleftharpoons \text{Cd}$	-0,40
$\text{Co}^{2+} + 2\text{e}^{-} \rightleftharpoons \text{Co}$	-0,28
$\text{Ni}^{2+} + 2\text{e}^{-} \rightleftharpoons \text{Ni}$	-0,27
$\text{Sn}^{2+} + 2\text{e}^{-} \rightleftharpoons \text{Sn}$	-0,14
$\text{Pb}^{2+} + 2\text{e}^{-} \rightleftharpoons \text{Pb}$	-0,13
$\text{Fe}^{3+} + 3\text{e}^{-} \rightleftharpoons \text{Fe}$	-0,06
$2\text{H}^{+} + 2\text{e}^{-} \rightleftharpoons \text{H}_2(\text{g})$	0,00
$\text{S} + 2\text{H}^{+} + 2\text{e}^{-} \rightleftharpoons \text{H}_2\text{S}(\text{g})$	+0,14
$\text{Sn}^{4+} + 2\text{e}^{-} \rightleftharpoons \text{Sn}^{2+}$	+0,15
$\text{Cu}^{2+} + \text{e}^{-} \rightleftharpoons \text{Cu}^{+}$	+0,16
$\text{SO}_4^{2-} + 4\text{H}^{+} + 2\text{e}^{-} \rightleftharpoons \text{SO}_2(\text{g}) + 2\text{H}_2\text{O}$	+0,17
$\text{Cu}^{2+} + 2\text{e}^{-} \rightleftharpoons \text{Cu}$	+0,34
$2\text{H}_2\text{O} + \text{O}_2 + 4\text{e}^{-} \rightleftharpoons 4\text{OH}^{-}$	+0,40
$\text{SO}_2 + 4\text{H}^{+} + 4\text{e}^{-} \rightleftharpoons \text{S} + 2\text{H}_2\text{O}$	+0,45
$\text{Cu}^{+} + \text{e}^{-} \rightleftharpoons \text{Cu}$	+0,52
$\text{I}_2 + 2\text{e}^{-} \rightleftharpoons 2\text{I}^{-}$	+0,54
$\text{O}_2(\text{g}) + 2\text{H}^{+} + 2\text{e}^{-} \rightleftharpoons \text{H}_2\text{O}_2$	+0,68
$\text{Fe}^{3+} + \text{e}^{-} \rightleftharpoons \text{Fe}^{2+}$	+0,77
$\text{NO}_3^{-} + 2\text{H}^{+} + \text{e}^{-} \rightleftharpoons \text{NO}_2(\text{g}) + \text{H}_2\text{O}$	+0,80
$\text{Ag}^{+} + \text{e}^{-} \rightleftharpoons \text{Ag}$	+0,80
$\text{Hg}^{2+} + 2\text{e}^{-} \rightleftharpoons \text{Hg}(\ell)$	+0,85
$\text{NO}_3^{-} + 4\text{H}^{+} + 3\text{e}^{-} \rightleftharpoons \text{NO}(\text{g}) + 2\text{H}_2\text{O}$	+0,96
$\text{Br}_2(\ell) + 2\text{e}^{-} \rightleftharpoons 2\text{Br}^{-}$	+1,07
$\text{Pt}^{2+} + 2\text{e}^{-} \rightleftharpoons \text{Pt}$	+1,20
$\text{MnO}_2 + 4\text{H}^{+} + 2\text{e}^{-} \rightleftharpoons \text{Mn}^{2+} + 2\text{H}_2\text{O}$	+1,23
$\text{O}_2(\text{g}) + 4\text{H}^{+} + 4\text{e}^{-} \rightleftharpoons 2\text{H}_2\text{O}$	+1,23
$\text{Cr}_2\text{O}_7^{2-} + 14\text{H}^{+} + 6\text{e}^{-} \rightleftharpoons 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$	+1,33
$\text{Cl}_2(\text{g}) + 2\text{e}^{-} \rightleftharpoons 2\text{Cl}^{-}$	+1,36
$\text{MnO}_4^{-} + 8\text{H}^{+} + 5\text{e}^{-} \rightleftharpoons \text{Mn}^{2+} + 4\text{H}_2\text{O}$	+1,51
$\text{H}_2\text{O}_2 + 2\text{H}^{+} + 2\text{e}^{-} \rightleftharpoons 2\text{H}_2\text{O}$	+1,77
$\text{Co}^{3+} + \text{e}^{-} \rightleftharpoons \text{Co}^{2+}$	+1,81
$\text{F}_2(\text{g}) + 2\text{e}^{-} \rightleftharpoons 2\text{F}^{-}$	+2,87

Increasing oxidising ability/*Toenemende oksiderende vermoë*Increasing reducing ability/*Toenemende reduserende vermoë*